An Advanced Materials Technology for Recovering Metals from Dilute Aqueous Streams

Jerome P. Downey, Ph.D., P.E. Professor of Extractive Metallurgy Campus Director, MUS Materials Science Ph.D. Program Montana Tech

September 20, 2017



Presentation Outline

- Introduction & Background
- Nanocomposite Particle Synthesis & Characteristics
- Continuous Flow Reactor Design & Operation
 - 3rd Generation
 - 4th Generation
- Metal Recovery via Electrowinning
- Present status/Concluding remarks
- Acknowledgements
- Questions

Montana Tech

Example application: metal-contaminated mine waters



Figure 1 – Selected metal concentrations of mine site samples

Background

- Dr. Arijit Bose (University of Rhode Island) developed a concept for a continuous magnetic field reactor for special ion exchange applications.
- U.S. Patent No. 6,635,181, Continuous Hybrid Magnetic Field Gradient Rotating Wall Device for Colloidal Magnetic Affinity Separations.
- URI and Hazen Research, Inc. (Golden, Colorado) collaborated on a Phase I STTR project to design, construct, and operate a bench-scale reactor.



Figure 2 – 1st generation reactor



Figure 3 – 2nd generation reactor

Metal capture efficiency

Particle retention



Figure 4 – Discharge solution analysis vs. time for the silver surrogate solution trial



Figure 5 – Photograph showing magnetic IX particles collected by the electromagnet

Nanocomposite particle synthesis

Magnetite core (20 – 30 nm)

Synthesis of core-shell nanoparticles using PAA as the metal-capturing ligand begins with tetrethyoxysilane (TEOS) hydrolysis.

Polyallylamine (PAA) was used as the metal-capturing ligand





Figure 6 – magnetic particles gathered with a permanent magnet following the TEOS step.

(performed in Dr. Rosenberg's lab at UM)

Nanocomposite particle characteristics

Energy dispersive X-ray measurements (EDX) gave a percent by iron of 80-81 % and a 10 % Si, consistent with deposition of a thin silica layer.

Dynamic Light Scattering (DLS) measurements gave a Z averaged diameter (ZAD) of 162 nm.

SEM and TEM images of the core-shell nanoparticles revealed particle aggregates of about 200 nm but individual 10-12 nm core-shell nanoparticles can be clearly identified.

Copper loading capacities of 0.5 mmole/gram.

The nanoparticle loading rates are about *10 times faster* than the related micro-scale silica polyamine composites.



P 44a No Bsa 1.tif Rosenberg P 44a No BSA Print Nag: 105000x 0 51 nn 9:30 03/18/16 TEM Mode: Imaging Microscopist: J Driver

100 nm HV=75kV Direct Nag: 200000x University of Montana

Figure 7 – TEM image of a cluster of nanocomposite particles

3rd generation continuous flow reactor

The design is much simpler and more robust than the previous prototypes:

- 100 liter total reactor volume
- Solution flow rates range from 2 to 10 liters per minute
- A single internal electromagnet replaces multiple external magnets
- Static mixers replace magnets intended to oscillate particles
- No moving parts except for pumps and solenoid valves



Figure 8 – 3rd generation continuous flow reactor system with in-line magnet module

System operation

A slurry of magnetic nanocomposite ion exchange (IX) particles is introduced to the surrogate wastewater stream at the top of the platform.

Functional groups impregnated on the surfaces of the ion exchange particles bond with metal ions as the particle-laden stream flows through the reactor.

The electromagnet captures and retains the nanocomposite particles without impeding the wastewater flow.

Valve positions are periodically changed to divert flow through a parallel magnetic module.

The electromagnet is de-energized to release the particles, which are subsequently stripped, reconditioned, and returned to service.

Figure 9 – The in-line electromagnet module consistently achieves magnetic particle capture efficiencies in excess of 98%.

System operation



Figure 10 – Schematic diagram depicts IX system operating sequence

4th generation continuous flow reactor

Figure 11 – the modified electromagnet configuration

Figure 12 – the 4th generation continuous flow reactor (120 L)

Metal recovery

Figure 13 – the laboratory EMEW[™] electrowinning circuit.

Figure 14 – copper and zinc sheets

System performance:

Copper

40-45 °C electrolyte 2.5V 11A 94% current efficiency

Zinc

30-50 °C electrolyte 5V 20A 95.5% current efficiency.

Present status

The culminating achievements of this research project include:

- 1. Successful development of a method for preparing magnetic nanocomposite IX particles that have proven effective in both laboratory and pilot plant evaluations
- 2. Two pilot-scale continuous flow reactor systems have been constructed, commissioned and successfully operated. These systems will be used to further develop and demonstrate the process.
- 3. Several Montana wastewater streams have been sampled and characterized; candidates for treatment via the magnetic nanocomposite/continuous flow reactor technology have been identified.

Concluding remarks

The combination of these innovative ideas represents a powerful technology for rapid continuous extraction of valuable and/or toxic metals from a broad range of contaminated waters

- continuous IX method proven capable of efficiently capturing metals and other contaminants from dilute aqueous streams
- mechanically very simple and can be configured for site-specific applications
- amenable to automated control systems
- cost effective low low capital and operating costs

Acknowledgements

The Montana University System and the MREDI program.

Montana Bureau of Mines & Geology

Past Collaborators & Coauthors:

- Dr. Arijit Bose, University of Rhode Island
- Dr. Guy Fredrickson, Idaho National Laboratory

Acknowledgements: Project Team

Montana Tech

Hsin Huang, Professor, Metallurgical and Materials Engineering Alysia Cox, Assistant Professor, Chemistry and Geochemistry Katherine Zodrow, Assistant Professor, Environmental Engineering David Hutchins, Materials Science Ph.D. Student Renee Schmidt, Geochemistry MS student Maureen Chorney, Metallurgical Engineering MS Student Jared Geer, Metallurgical & Materials Engineering undergraduate student Auva Speiser, Metallurgical & Materials Engineering undergraduate student Elizabeth Raiha, Metallurgical & Materials Engineering undergraduate student Christina Eggensperger, Environmental Engineering post-bac student

University of Montana

Edward Rosenberg (Co-PI), Professor, Department of Chemistry & Biochemistry Ryan Latterman, Post-Doctoral Research Associate Emil DeLuca, Research Associate Lab Manager

MontanaTech

References

- 1. Y. Huang *Water Research* 80 (2015) 159-168
- 2. M. Ma Colloids and Surfaces A: Physicochem. Eng. Aspects 212 (2003) 219-226
- 3. Y. Xia et. al. *Nano Letters* (2002) 2, 183-186
- 4. Helfferich, F., Ion Exchange, McGraw-Hill Series in Advanced Chemistry, McGraw-Hill Book Company, New York, 1962.
- 5. G. Cifuentes, Model and Simulation of an Ion Exchange Process for Extraction of Antimony," Collected Proceedings of the 140th Annual Meeting & Exhibition, TMS, 2011.
- 6. Rosenberg, E, Characterization of Surface-Bond Zr(IV) and Its Application to Removal of As(V) and As(III) from Aqueous System Using Phosphonic Acid Modified Nanoporous Silica Polyamine Composites, *Ind. Eng. Chem. Res.*, **2009**, 48(8), pp 3991-4001.
- 7. W. Stumm, Chemistry of the Solid-Water Interface A Wiley-Interscience Publication, New York, 1992.
- 8. G. Tochobanoglous et al., Wastewater Engineering, Treatment, Disposal and Reuse, McGraw-Hill, New York, 1991
- 9. H.C. Thomas, Heterogeneous ion exchange in a flowing system, J. Am. Chem. Soc. 66, 1664-1666 (1944)
- Z. Yu, et al., Application of mathematical models for ion-exchange removal of calcium ions from potassium chromate solutions by Amberlite IRC 748 resin in a continuous fixed bed column, Hydrometallurgy, 158 165-171 (2015)